

Name: _____

How to put **ELECTRONS** in the rings of an ATOM...

Step 1: Find the **element** in the **periodic table**

Periodic Table of the Elements

1	2	3	4	5	6	7	8	9	10	11	12	13	14	15	16	17	18	19	20	21	22	23	24	25	26	27	28	29	30	31	32	33	34	35	36	37	38	39	40	41	42	43	44	45	46	47	48	49	50	51	52	53	54	55	56	57	58	59	60	61	62	63	64	65	66	67	68	69	70	71	72	73	74	75	76	77	78	79	80	81	82	83	84	85	86	87	88	89	90	91	92	93	94	95	96	97	98	99	100
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Step 2: Find the **atomic number** of the **element**

16	8
O	
Oxygen	
16.00	(6)

Step 3: NUMBER of PROTONS = NUMBER of ELECTRONS

Step 4: Fill the rings (circles around the nucleus) with the ELECTRONS.

See "RING RULES" below:

Ring # 1 = NO MORE THAN 2 ELECTRONS

(if RING # 1 is full, start filling RING # 2)

Ring # 2 = NO MORE THAN 8 ELECTRONS

(if RING # 2 is full, start filling RING # 3)



Ring # 3 = NO MORE THAN 8 ELECTRONS

